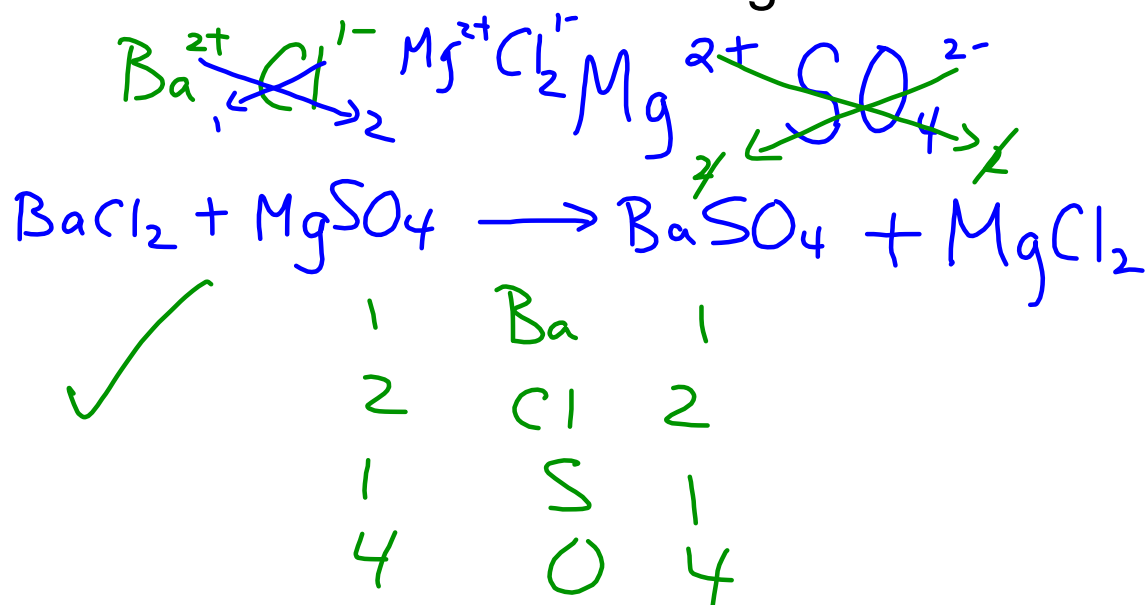
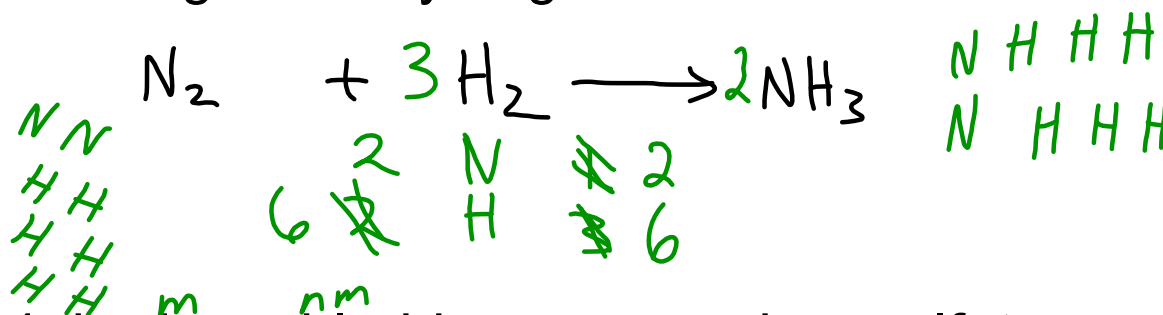
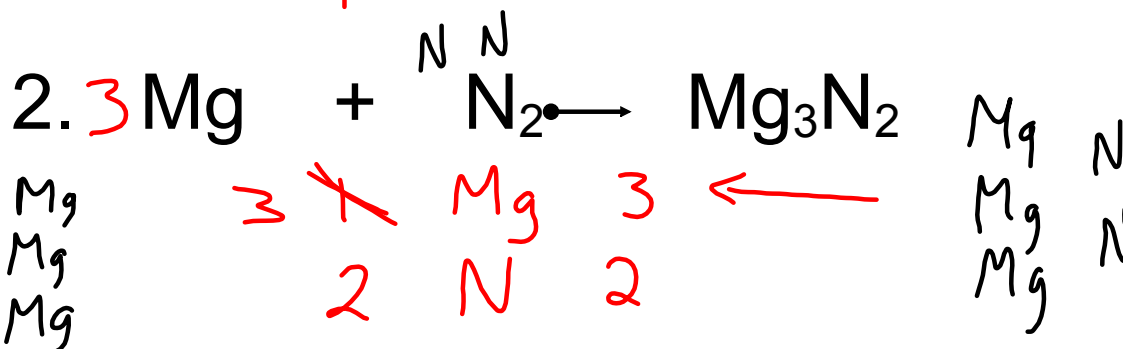
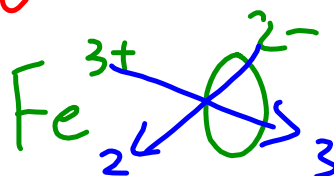
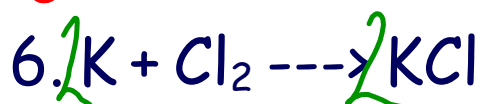
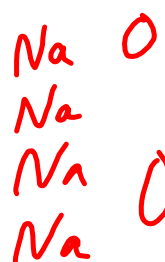
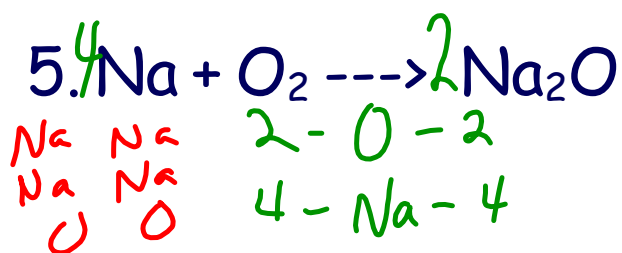
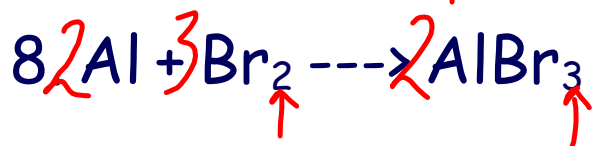
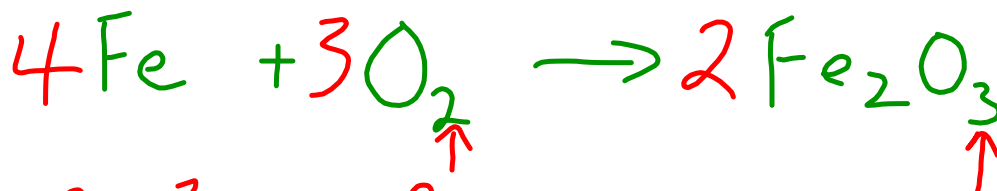


- Copy the following into your notebook & balance each:





7. iron + oxygen \rightarrow iron(III) oxide



$$\frac{1}{2} + \frac{1}{3} =$$

Practice: Review

Identify each as either being an ionic (I), polyatomic (P) or molecular (M) compound.

1. FeO = _____

4. NaOH = _____

2. K_3PO_4 = _____

5. Na_2S = _____

3. CH_4 = _____

6. H_2 = _____

7. carbon disulfide = ____ 10. calcium carbonate = ____

8. magnesium oxide = ____ 11. sulfur trioxide = ____

9. copper (II) sulfate = ____ 12. sulfuric acid = ____

Word & Skeletal Equations

- Always written as:



SOLID (s): A precipitate is formed

LIQUID (l):

GAS (g):

AQUEOUS (aq): The substance is dissolved in water or it is in solution form

Question

What are some ways in which we know a chemical reaction has occurred?

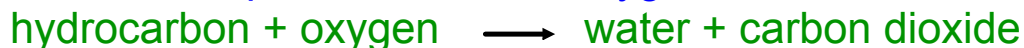
1. A color change
2. Formation of a gas
3. Formation of a precipitate
(insoluble solution falls to the bottom)
4. Energy change
(heat given off = exothermic; heat absorbed = endothermic)
5. Odor Change



Types of Chemical Equations Name: _____

Find examples for each on p 230,234 and 240-241

1. Combustion - rapid reaction with oxygen



Ex:

2. Synthesis - combine two small molecules (or elements) to make one larger molecule.



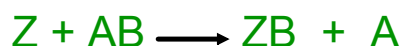
Ex:

3. Decomposition - splitting one large molecule into 2 or more smaller molecules or elements.



Ex:

4. Single Displacement - one element replaces another element from a compound.



Ex:

5. Double Displacement - elements from each compound are exchanged (replaced).



Ex:

What Type of Reactions are These?

